Running title: Leaf trypsin inhibitor and its peptides with antioxidant activities

# A novel trypsin inhibitor from sweet potato (*Ipomoea batatas Lam.*) leaves and its synthesized peptides with antioxidant activities *in vitro*

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## Abstract

Recombinant SPLTI-a [sweet potato leaf trypsin inhibitor-a] overproduced in *E. coli* (M15) was purified by $Ni^{2+}$-chelated affinity chromatography. The molecular mass of SPLTI-a is ca. 8000 Da as determined by sodium dodecyl sulfate-polyacrylamide gel electrophoresis (SDS-PAGE). SPLTI-a was examined using different antioxidative models (Total antioxidant status, reducing power method, $Fe^{2+}$-chelating ability, ferric thiocyanate (FTC) method, and protecting calf thymus DNA against hydroxyl radical-induced damage). The SPLTI-a protein with a concentration of 100 μg/mL exhibited highest activity (expressed as 2.12 ± 0.02 mM Trolox equivalent antioxidative value, TEAC) in total antioxidant status test. Like total antioxidant status, the reducing power, $Fe^{2+}$-chelating ability, FTC activity and protecting calf thymus DNA against hydroxyl radical-induced damage all showed that SPLTI-a polypeptide has significant antioxidant activities. It was found that the antioxidant activity increased after 24 h hydrolysis of SPLTI-a by trypsin from 18% (0 h) to about 35% (24 h). Accumulation of shorter peptides increased along the longer trypsin incubation. The obtained VR, STIEK, ITDGK, and EYIFDR showed $IC_{50}$ (concentration for 50% inhibition) values of 5.83, 3.75, 2.65, and 0.73 mM, respectively, when scavenging activity of DPPH radicals (%) was measured. These findings mean that tyrosine residue is most important in antiradical activities. It was suggested that SPLTI-a possess antioxidant activities.



**Keywords:** Antioxidant Activity; Leaf Trypsin Inhibitor; Recombinant Protein; Sweet Potato.

## INTRODUCTION

Plant proteinases inhibitors (PIs) are widely distributed in plants. Serine, cysteine and aspartic proteinases are the main proteinases inhibited. Most PIs have four conserved Cys residues forming two disulfide bonds in a single or double chain polypeptide (Oliva, et al., 2010). In higher plants, PIs are shown to be particularly abundant in storage organs, such as tubers (Richardson, 1991). A large body of evidence indicated that PIs functioned as storage proteins, regulate the endogenous proteinase activities (Dunaevsky, et al., 1998), and suppress the exogenous proteinase activities from pathogens and pests (Ryan, 1989). Four types of PIs were also found to accumulate rapidly in leaves in response to mechanical wounding or insect chewing (Constabel, 1999), suggesting a direct role of PIs in plant protection.

Reactive oxygen species (ROS) can be formed by both endogenous and exogenous sources in living organisms. Within the cells, ROS can be generated in mitochondrial and microsomal electron transport systems, in soluble oxidase enzyme systems, and during phagocyte activation (Kehrer, 1993). Exogenous ROS generate from air and water pollutants, cigarette smoke, organic solvents, heavy metals, certain drugs, and radiation, etc (Valko et al., 2006). The ROS play an important role related to the degenerative or pathological processes of various serious diseases, such as aging, cancer, coronary heart disease, Alzheimer's disease, neurodegenerative disorders, atherosclerosis, cataracts, and inflammation (Huang, et al., 2010). The use of food is widespread and plants still present a large source of natural antioxidants that might serve as leads for the development of novel drugs. Several antiinflammatory, antinecrotic, neuroprotective and hepatoprotective drugs have recently been shown to have an antioxidant and/or antiradical scavenging mechanism (Repetto and Llesuy, 2002). During the last few years, natural antioxidants and

compounds with radical scavenging activity have been found, such as phenolic compounds (Huang et al., 2008), anthocyanin (Espin et al., 2000), water extract of *Flemingia* species (Hsieh et al., 2010), thioredoxin *h* protein (Huang et al., 2004a), sporamin (Huang et al., 2007), and mucilage (Huang et al., 2006) from sweet potato root. The objectives of this work were to investigate the antioxidant property of SPLTI-a from sweet potato leaf in comparison with chemical compounds such as butylated hydroxytoluene (BHT), reduced glutathione or ascorbate in a series of *in vitro* tests.

## MATERIALS AND METHODS

### Expression of SPLTI-a in *E. coli*

*SPLTI-a* (Gene Bank accession number: AF330700) was expressed in *E. coli*. The coding sequence was amplified from *SPLTI-a* cDNA using an oligonucleotide (5´-GGA TCC AGA AAA TGC AGC GCA TCA C -3´), with a *Bam*HI site (underlined) at the putative initial Met residue, and an oligonucleotide (5´-AGAAC TCCGTC GAT AAG CTT GGT -3´), with a *Hind*III site at the 3´ end. The PCR fragment was subcloned in pGEM T-easy vector. And the plasmid was then digested with *Bam*HI and *Hind*III and subcloned in pQE32 expression vector (QIAexpress expression system, Qiagen). The resulting plasmid, termed pQE-SPLTI-a, was introduced into *E. coli* (M15). Cultures of the transformed *E. coli* (M15) overexpressed a protein of the expected molecular mass, which was purified by affinity chromatography in Ni-nitrilotriacetic acid (NTA) columns (Qiagen), according to the manufacturer's instructions.

Protein Staining on 15% SDS-PAGE Gels

SPLTI-a was detected on 15% SDS-PAGE gels. Samples treated with sample buffer and β-mercaptoethanol (2-ME) with a final concentration of 14.4 mM were heated at $100^0$C for 5 min before 15% SDS-PAGE.

Measurement of Total Antioxidant Status

Total antioxidant status of the SPLTI-a protein was measured using the total antioxidant status assay kit (Calbiochem Corp) according to the manufacturer's instructions. The assay relies on the antioxidant ability of the protein to inhibit oxidation of 2, 2' azino-bis-[3-ethylbenz-thiazoline-6-sulfonic acid] (ABTS) to ABTS*$^+$ by metmyoglobin. The amount of ABTS*$^+$ produced is monitored by reading the absorbance at 600 nm. Under these reaction conditions, the antioxidant ability of SPLTI-a protein decreases the absorbance at 600 nm in proportion to its concentration. The final antioxidant capacity of SPLTI-a protein was calculated by the following formula: Trolox equivalent value (mM) = [factor x (absorbance of blank-absorbance of sample)]; factor= [concentration of standard/(absorbance of blank-absorbance of standard)].

Scavenging Activity against DPPH Radical

DPPH is a relatively stable free radical which when encounters proton donors such as antioxidants, the radicals get quenched and absorbance gets reduced. The effect of SPLTI-a on the DPPH radical was estimated according to the method of Huang et al (2004b). An aliquot of SPLTI-a (30 μL) was mixed with 100 mM Tris-HCl buffer (120 μL, pH 7.4) and then 150 μL of the DPPH in ethanol with a final concentration of 250 μM was added. The mixture was shaken vigorously and left to stand at room temperature for 20 min in the dark. The absorbance at 517 nm of the

reaction solution was measured spectrophometrically. The percentage of DPPH decolourization of the sample was calculated according to the equation: % decolourization= 【 1- $Abs_{sample}$ /$Abs_{control}$ 】×100. The $IC_{50}$ values denote the concentration of sample which is required to scavenge 50% of DPPH free radicals.

Determination of Antioxidant Activity by Reducing Power Measurement

Reducing power method indirectly evaluates the antioxidant activity. The reducing powers of the SPLTI-a and glutathione were determined according to the method of Chang et al. (2007). SPLTI-a (0, 0.2, 0.4, 0.8, 1.0, and 1.2 mg/mL) or glutathione was mixed with an equal volume of 0.2 M phosphate buffer, pH 6.6, and 1% potassium ferricyanide. The mixture was incubated at 50 $^{0}$C for 20 min, during which time ferricyanide was reduced to ferrocyanide. Then an equal volume of 1% trichloroacetic acid was added to the mixture, which was then centrifuged at 6,000 *g* for 10 min. The upper layer of the solution was mixed with deionized water and 0.1% $FeCl_3$ at a radio of 1 : 1 : 2, and the absorbance at 700 nm was measured to determine the amount of ferric ferrocyanide (Prussian Blue) formed. Increased absorbance of the reaction mixture indicated increased reducing power of the sample.

Determination of Antioxidant Activity by $Fe^{2+}$-Chelating Ability

The $Fe^{2+}$-chelating ability was determined according to the method of Huang et al. (2007). The $Fe^{2+}$ was monitored by measuring the formation of ferrous iron-ferrozine complex at 562 nm. SPLTI-a (0, 0.25, 0.5, 1, 2, and 4 mg/mL) was mixed with 2 mM $FeCl_2$ and 5 mM ferrozine at a ratio of 10 : 1 : 2. The mixture was shaken and left to stand at room temperature for 10 min. The absorbance of the resulting solution at 562 nm was measured. The lower the absorbance of the reaction

mixture the higher the $Fe^{2+}$-chelating ability. The capability of the sample to chelate the ferrous iron was calculated using the following equation:

Scavenging effect (%) = 【 1- $Abs_{sample}$ / $Abs_{control}$ 】×100

Determination of Antioxidant Activity by the Ferric Thiocyanate (FTC) Method

The FTC method was adapted from the method of Osawa and Namiki (1981). Twenty mg/mL of samples dissolved in 4 ml of 99.5% (*w/v*) ethanol were mixed with linoleic acid (2.51%, *v/v*) in 99.5% (*w/v*) ethanol (4.1 mL), 0.05 M phosphate buffer pH 7.0 (8 mL) and deionized water (3.9 mL) and kept in a screw-cap container at 40℃ in the dark. Then, to 0.1 mL of this solution was added 9.7 mL of 75% (*v/v*) ethanol and 0.1 mL of 30% (*w/v*) ammonium thiocyanate. Precisely 3 min after the addition of 0.1 mL of 20 mM ferrous chloride in 3.5% (*v/v*) hydrochloric acid to the reaction mixture, the absorbance at 500 nm of the resulting red color [$Fe(SCN)^{2+}$, $Fe^{3+}$ was formed after linoleic acid peroxide was produced and Fenton reaction occurred.] was measured every 24 h until the day when the absorbance of the control reached the maximum value. The inhibition of linoleic acid peroxidation was calculated as (%) inhibition = 100 - [(absorbance increase of the sample/absorbance increase of the control) x 100]. All tests were run in duplicate and analyses of all samples were run in triplicate and averaged.

Protection of SPLTI-a against Hydroxyl Radical-Induced Calf Thymus DNA Damage

The hydroxyl radical was generated by Fenton reaction according to the method of Kohno et al. (1991). The 15 μL reaction mixture containing SPLTI-a (1.25, 2.5, 5, and 10 mg/mL), 5 μL of calf thymus DNA (1 mg/mL), 18 mM $FeSO_4$, and 60 mM hydrogen peroxide were incubated at room temperature for 15 min. Then 2 μL of 1

mM EDTA was added to stop the reaction. Blank test contained only calf thymus DNA and the control test contained all reaction components except SPLTI-a. The treated DNA solutions were subjected to agarose electrophoresis and then stained with ethidium bromide and examined under UV light.

## Determination of the Antioxidative Activity of SPLTI-a Tryptic Hydrolysates

Six mg of SPLTI-a was dissolved in 1 mL of 0.1 M KCl buffer (pH 8.0). Then 0.1 mL (12 mg) of trypsin was added at 37 $^{0}$C for 0 and 24 h. After hydrolysis, 0.5 mL of 0.5 M Tris-HCl buffer (pH 8.3) was added, and the solution was heated at 100 $^{0}$C for 5 min to stop enzyme reaction. The trypsin was heated before SPLTI-a hydrolysis for the 0 h reaction. Each of the 60 μL SPLTI-a hydrolysates was used for determinations of the DPPH antioxidative activities by spectrophotometry (Mine et al., 2004; Qian et al., 2008).

## Chromatograms of Tryptic Hydrolysates of SPLTI-a on a Sephadex G-50 Column

The unhydrolyzed SPLTI-a and tryptic SPLTI-a hydrolysates at 24 h were separated by Sephadex G-50 chromatography (1 x 60 cm). The column was eluted with 20 mM Tris-HCl buffer (pH 7.9). The flow rate was 30 mL/h, and each fraction contained 2 mL of which the absorbance at 280 nm was then determined.

## Statistical Analysis

Means of triplicates were calculated. Student's *t* test was used for comparison

between two treatments. All data (expressed as percent of control value) were means± SE. A difference was considered to be statistically significant when $p < 0.05$, $p < 0.01$ or $p < 0.001$.

## RESULTS and DISCUSSION

### Purification of expressed SPLTI-a

*SPLTI-a* cDNA clones from sweet potato leaf was isolated. *SPLTI-a* was subcloned in a pQE-32 expression vector in *E. coli* and SPLTI-a was produced with a 6x His-tag at the N-terminus. SDS-PAGE analysis of crude extracts from transformed *E. coli* (M15) showed a high level of a polypeptide with the expected molecular mass (ca. 8 kDa). This polypeptide was found as a soluble protein in the supernatant (Fig. 1, lane 2), and was absent in protein extracts obtained from *E. coli* transformed with pQE-32 vector (Fig. 1, lane 1). The expressed protein was purified from crude extracts by $Ni^{2+}$ -chelate affinity chromatography, which yielded highly purified His-tagged SPLTI-a (Fig. 1, lane 3).

### Measurement of Total Antioxidant Status using ABTS assay

Several methods have been developed to determine the antioxidant potential of natural products. The trolox equivalent antioxidant capacity (TEAC) using ABTS as an oxidant, the ferric reducing antioxidant power (FRAP), and the DPPH free radical

scavenging assays are some of the most commonly used. Antioxidants can reduce radicals primarily by two mechanisms: single electron transfer and hydrogen atom transfer. ABTS, FRAP, and DPPH are methods that measure the former (Ozgen, et al., 2006).

ABTS assay is often used in evaluating total antioxidant power of single compounds and complex mixtures of various plants. It is used to estimate the total antioxidant power because the method is quick and simple to perform, and the reaction is reproducible and linearly related to the molar concentration of the antioxidants (Benzie et al., 1999). This was measured using the total antioxidant status assay kit (Fig. 2A). SPLTI-a protein exhibited a dose-dependent total antioxidant activity within the applied concentrations (0, 2.5, 5, 10, 20, 40, 60, 80, and 100 µg/mL), the highest at 100 µg/mL (expressed as 2.12 ± 0.02 mM Trolox equivalent antioxidative value, TEAC). At 2.5 µg/mL, SPLTI-a displayed the lowest total antioxidant status (1.09 ± 0.01 mM TEAC).

Measurement of Reducing Power

In this assay, the yellow color of the test solution changes to various shades of green and blue, depending on the reducing power of each compound. Presence of reducers causes the conversion of the $Fe^{3+}$/ferricyanide complex used in this method to the ferrous form. By measuring the formation of Perl's Prussian blue at 700 nm, it is possible to determine the $Fe^{2+}$ concentration (Gülçin et al., 2003). We investigated

244 the $Fe^{3+}$-$Fe^{2+}$ transformation in the presence of the samples of SPLTI-a to measure its
245 reducing capacity. The reducing capacity of a compound may serve as a significant
246 indicator of its potential antioxidant activity (Meir et al., 1995). The antioxidant
247 activity of putative antioxidants have been attributed to various mechanisms, among
248 them are prevention of chain initiation, binding of transition metal ion catalysts,
249 decomposition of peroxides, prevention of continued hydrogen abstraction, and
250 radical scavenging (Liu et al., 2006). The reducing power of SPLTI-a is shown in Fig.
251 2B with glutathione served as a positive control. The reducing power activity of
252 SPLTI-a exhibited a dose-dependence (significant at $p< 0.05$) within the applied
253 concentrations (0, 0.1, 0.2, 0.4, 0.8, 1.0, and 1.2 mg/mL).
254
255 Measure of $Fe^{2+}$-chelating ability

256 Direct reaction of a substance is not the only mechanism by which the
257 antioxidants may display their activity. Antioxidants act through numerous possible
258 mechanisms. Some antioxidants do not convert free radicals to more stable products
259 but slow the rate of oxidation by several different mechanisms. One of the most
260 important mechanisms of action of secondary antioxidants is chelation of pro-oxidant
261 metals. Iron and other transition metals (copper, chromium, cobalt, vanadium,
262 cadmium, arsenic, nickel) promote oxidation by acting as catalysts of free radical
263 reactions. These redox-active transition metals transfer single electrons during

changes in oxidation states. Chelation of metals by certain compounds decreases their pro-oxidant effect by reducing their redox potentials and stabilizing the oxidized form of the metal. Chelating compounds may also sterically hinder formation of the metal hydroperoxide complex (Blokhina et al 2003).

$Fe^{2+}$ ion is the most powerful pro-oxidant among the various species of metal ions. Ferrozine can quantitatively form complexes with $Fe^{2+}$. Therefore, the measurement of color reduction allowed the estimation of the metal chelating activity of the coexisting chelator. Metal chelating capacity was significant because it reduced the concentration of the catalyzing transition metal in lipid peroxidation (Halliwell and Gutteridge, 1984). The metal chelating capacity of SPLTI-a and standard antioxidants was determined by assessing their ability to compete with ferrozine for the ferrous ions. The $Fe^{2+}$-chelating ability of the SPLTI-a is shown in Fig. 2C. EDTA was used as a positive control. The $Fe^{2+}$-chelating ability of SPLTI-a was lower than that of EDTA and this difference was statistically significant ($P < 0.05$). SPLTI-a at doses of 0.25, 0.5, 1, 2, and 4 mg/mL exhibited 42.36, 44.71, 55.94, 62.31, and 68.02% iron binding capacity, respectively. On the other hand, EDTA at doses of 0.05, 0.1, 0.2, 0.4, and 0.8 mg/mL had 37.24, 67.59, 78.12, 84.25 and 87.68 % chelating activity of iron, respectively.

Ferric thiocyanate (FTC) method

Malondialdehyde formed from the breakdown of polyunsaturated fatty acids was served as a convenient index for determining the extent of lipid perxoidation reaction. The ferric thiocyanate method measured the amount of peroxide produced during the initial stages of oxidation which was the primary product of oxidation. So far

low-density lipoprotein (LDL) peroxidation has been reported to contribute to atherosclerosis development (Steinbrecher, 1987). Therefore, delay or prevention of LDL peroxidation is an important function of antioxidants. Fig. 2D shows the time-course curve for the antioxidative activity of the SPLTI-a from sweet potato leaf, BHT, and $H_2O$ by the FTC method. The BHT was used as a positive control, and $H_2O$ as a negative control. The results indicate that SPLTI-a has antioxidative activity. SPLTI-a may act as a significant LDL peroxidation inhibitor ($P < 0.05$).

Protection against Hydroxyl Radical-Induced Calf Thymus DNA Damage by SPLTI-a

Free radicals could damage macromolecules in cells, such as DNA, protein, and lipids in membranes (Kohno et al., 1991). The oxidative damage of DNA is one of the most important mechanisms in the initiation of cancer. The damage is usually caused by hydroxyl radicals. The activity of these radicals can be reduced by natural antioxidants found in plants including herbs. The Fenton reaction involves the reaction between hydrogen peroxide and $Fe^{2+}$ to form hydroxyl radicals. Scavengers of hydroxyl radicals inhibit this reaction through the reduction of $Fe^{2+}$. Fig. 3 shows that SPLTI-a protected calf thymus DNA against hydroxyl radical-induced damages. The blank contained calf thymus DNA only, and the control contained all components except SPLTI-a. Compared to the blank and control, it was found that 2.5 mg/mL SPLTI-a could protect against hydroxyl radical induced calf thymus DNA damages during 15-min reactions.

## Determination of the Antioxidative Activity of tryptic SPLTI-a Hydrolysates and Their Peptide Distributions

We used trypsin to hydrolyze SPLTI-a to mimic the hydrolysis course during digestion in human's (or animal's) intestine. Fig. 4 shows the antioxidative activity of tryptic SPLTI-a hydrolysates and the antioxidative activity (scavenging activity of DPPH radicals, percent) of tryptic SPLTI-a hydrolysates collected at different trypsin hydrolysis times. From the results, it was found that the antioxidative activity increased from 18% (0 h) to about 35% (24 h); at the same time, smaller peptides increased with trypsin hydrolytic time. Here, 100% activity was defined as the concentration of substrate that causes 100% loss of the DPPH activity. The purifications of potential peptides of antioxidative activity need further investigations. We used synthetic peptides to measure antioxidative activity. Synthetic peptides were designed by mimicking trypsin cutting sites of SPLTI-a gene products from sweet potato (http://expasy.nhri.org.tw/tools/peptidecutter/). New peptides (Table I) for antioxidative activity, that is, VR, STIEK, ITDGK and EYIFDR were synthesized according to deduction. $IC_{50}$ values of individual peptides were 5.83, 3.75, 2.65, and 0.73 mM, respectively, when scavenging activity of DPPH radicals (%) was measured. Tyrosine residues with free –OH were reported to have antioxidant activities. Tyrosine had an effect on DPPH radical scavenging, ABTS radical scavenging, superoxide anion radical scavenging, $H_2O_2$ scavenging, total ferric ions reducing power and metal chelating on ferrous ions activities (Gu¨lcin, 2007). It was reported that the radical scavenging and antioxidant activities of tyrosine were highly controlled by the number of phenolic hydroxyl groups. Tyrosine residue could be oxidized to a tyrosyl radical through an electron-transfer process. Nitration may

follow from the reaction of a tyrosyl radical with ·NO2 (Lin, et al 2003; Pietraforte, et al., 2001). Our results further indicate that tyrosine residues (EYIFDR) in sweet potato SPLTI-a contribute to the antiradical activity. The synthetic peptide, EYIFDR, has the highest antioxidant activity ($IC_{50}$ is 0.73 mM) as good as reduced glutathione ($IC_{50}$ is 0.07 mM). These results demonstrated that deduced synthetic peptides from tryptic SPLTI-a hydrolysates exhibited antioxidative activity.

Especially, EYIFDR was more potent for scavenging hydroxyl and peroxyl radicals. It is generally accepted that the chemical activity of hydroxyl radical is the strongest among ROS. Hydroxyl radical reacts easily with biomolecules, such as amino acids, proteins and DNA. Therefore, the removal of hydroxyl radical is probably one of the most effective defenses of a living body against various diseases (Cacciuttolo et al., 1993). Carbon-centered radicals that represent R · , RO · and ROO · could be quenched by EYIFDR. This result agreed with a lipid peroxidation inhibition assay finding that EYIFDR inhibited lipid peroxidation by scavenging lipid-derived radicals.

In conclusion, the results from *in vitro* experiments, including total antioxidant status assay (Fig. 2A), reducing power method (Fig. 2B), $Fe^{2+}$-chelating ability (Fig. 2C), FTC method (Fig. 2D), and hydroxyl radical-induced calf thymus DNA damage (Fig. 3), demonstrated that SPLTI-a in sweet potato leaf may have significant antioxidant activities. SPLTI-a may contribute significantly to change the redox states and as a potent antioxidant against both hydroxyl and peroxyl radicals when people consume sweet potato. The *ex vivo* or *in vivo* antioxidant activity of SPLTI-a should

be performed in near further.


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# Figure Legends

Figure 1. SDS-PAGE analysis of purified recombinant sweet potato leaf trypsin inhibitor (SPLTI-a). Crude extracts from *E. coli* (M15) transformed with pQE30 (lane 1) or with pQE30- SPLTI-a (lane 2) were analyzed by 15% (*w/v*) SDS/PAGE with 10 μg protein applied on each lane, and then the gel was stained with Coomassie blue G-250. Molecular masses of standard proteins are indicated at the left of the figure. His-tagged SPLTI-a was purified by $Ni^{2+}$-chelated affinity chromatography (lane 3). The experiments were done twice and a representative one is shown.

Figure 2. Radical scavenging activity of the recombinant SPLTI-a determined by TEAC (A), reducing power (B), $Fe^{2+}$-chelating ability (C) and inhibition of linoleic acid peroxidation (D). Each absorbance value represents average of triplicates of different samples analyzed. Results represent the means ± SE from at least 3 separate experiments. $*p < 0.05$, $**p < 0.01$ and $***p < 0.001$ (unpaired t test) compared to SPLTI-a unsupplemented samples.

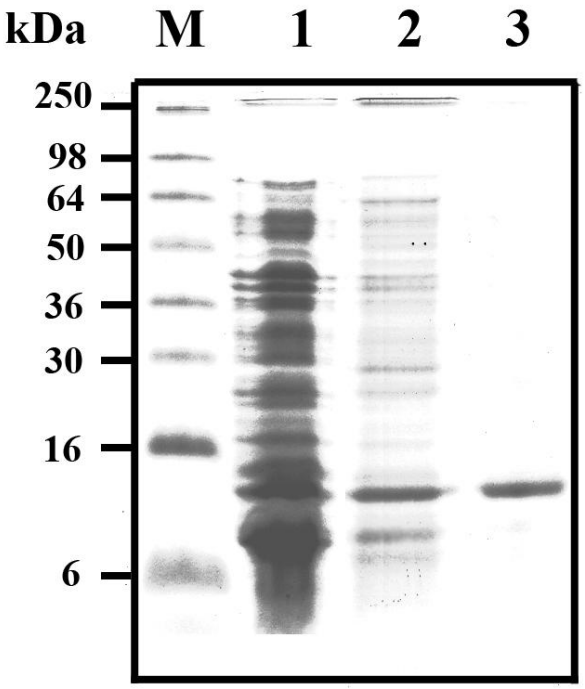
Figure 3. Protection against hydroxyl radical-induced calf thymus DNA damage by

recombinant SPLTI-a. Sample lanes 1-4 contained 1.25, 2.5, 5, and 10 mg/mL SPLTI-a, respectively. Blank (B) contained calf thymus DNA only; while the control (C) contained all reaction components except SPLTI-a.

Figure 4. Antioxidative activity of tryptic hydrolysates of recombinant SPLTI-a. The plot shows the antioxidative activity (%) of SPLTI-a hydrolysates at different trypsin hydrolysis time (0 hr and 24 hr). The proteins and the scavenging activity of DPPH radicals (%) were shown. The scavenging effect (%) was calculated according to the equation [1-(Abs 517 nm of sample ÷ Abs 517 nm of control)] × 100 %.

500

501 **Figure 1**

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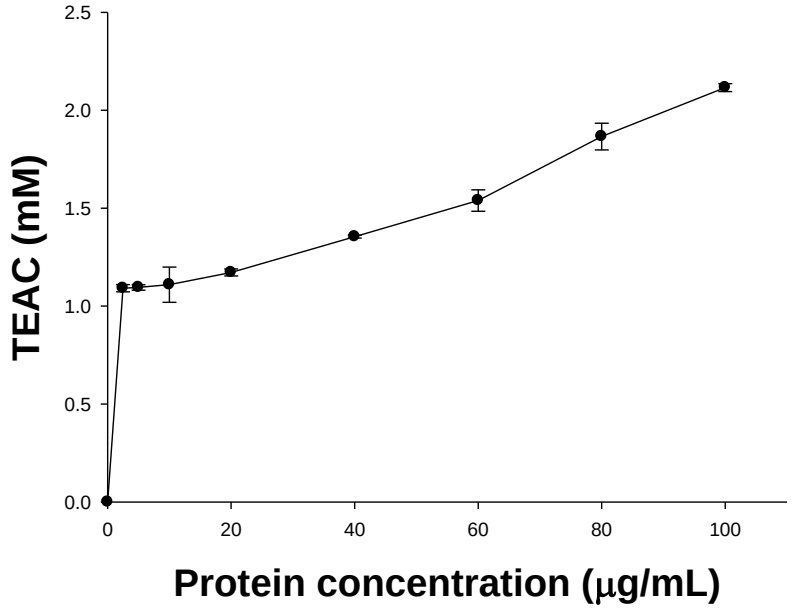
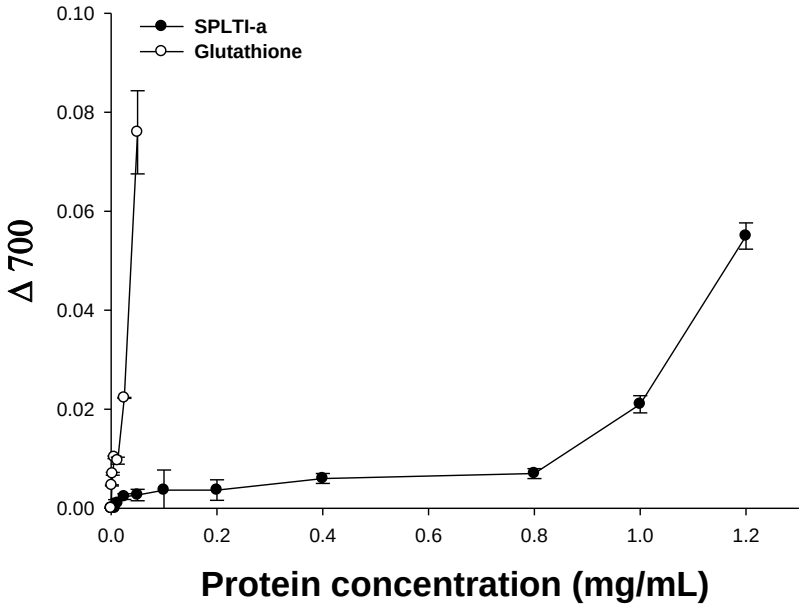
# Figure 2

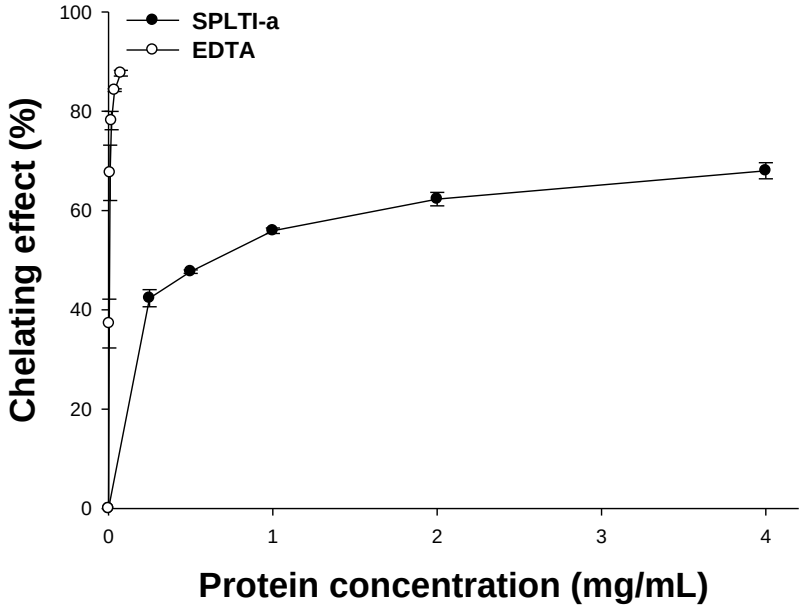
**A.**


TEAC (mM)
Protein concentration (μg/mL)
0.0
0.5
1.0
1.5
2.0
2.5
0
20
40
60
80
100

**B.**


SPLTI-a
Glutathione
Δ 700
Protein concentration (mg/mL)
0.00
0.02
0.04
0.06
0.08
0.10
0.0
0.2
0.4
0.6
0.8
1.0
1.2

512 **C.**

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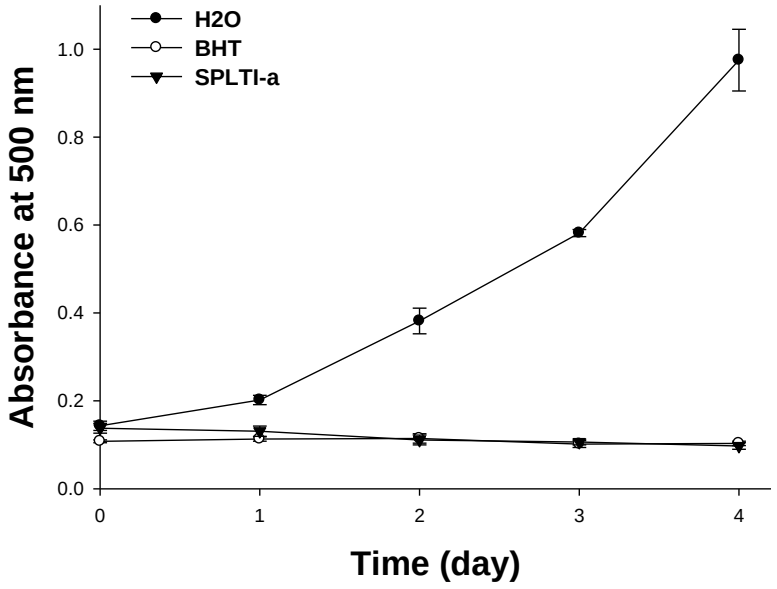
514 **D.**

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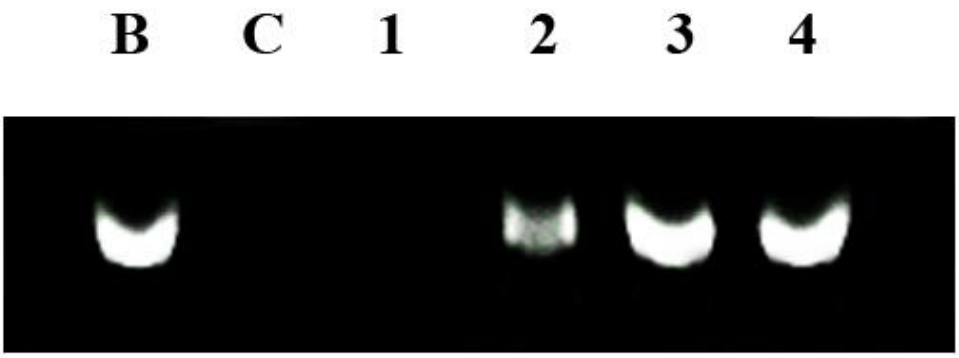
519 **Figure 3**

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527 **Figure 4**

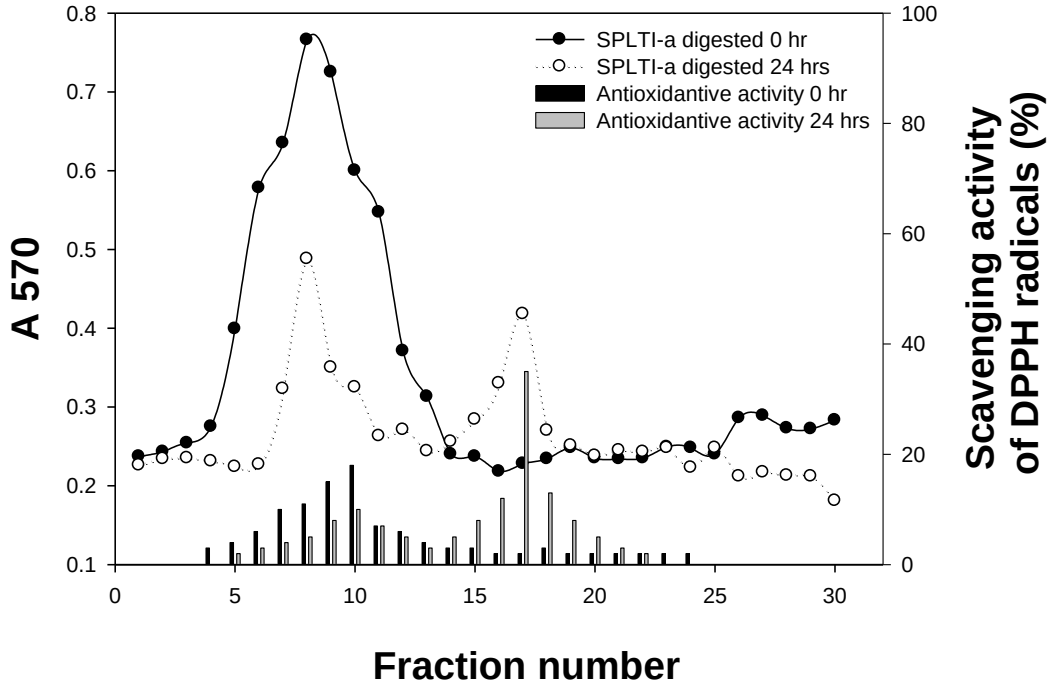


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Table 1: SPLTI-a peptides with antioxidant activity

| Peptides | Scavenging activity of DPPH radicals (%), $IC_{50}$ (mM) |
|---|---|
| GSH (control) | 0.07 ± 0.01 |
| VR | 5.83 ± 0.57 |
| STIEK | 3.75 ± 0.35 |
| ITDGK | 2.65± 0.24 |
| EYIFDR | 0.73 ± 0.02 |

Note: The sequence of SPLTI-a contains pre-pro-sequence. This sequence was retrieved from the NCBI (National Center for Biotechnology Information, http://www.ncbi.nlm.nih.gov) with the accession number AF330700.

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# 甘薯葉中新型胰蛋白酶抑制因子和其合成之胜肽含有抗氧化的活性

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在大腸桿菌(M15)中大量表現重組蛋白質 SPLTI-a，然後利用鎳離子螯合之親和性管柱純化。SPLTI-a 經 SDS-PAGE 分析其分子量約為 8 kDa. 本研究利用不同的抗氧化方法評估(總抗氧化能力、還原力、亞鐵離子螯合能力、抑制過氧化物形成能力, 和保護 DNA 免於氫氧自由基傷害)。SPLTI-a 在總抗氧化能力分析上在 100 μg/mL 時可達最高的抗氧化活性(以 2.12 ± 0.02 mM Trolox equivalent antioxidative value, TEAC, 表示)。 在所有分析項目中，重組之 SPLTI-a 蛋白質都具有顯著的的抗氧化活性。利用胰蛋白酶水解 SPLTI-a 時，小分子的胜肽會隨著水解時間增加。 24 小時後抗氧化活性(對 DPPH 之清除能力)可以從 18 % (0 h) 增加到 35 % (24 h)。利用電腦模擬胰蛋白酶水解 SPLTI-a 蛋白質的結果，四種人

562 工合成具有抗氧化活性胜肽: VR, STIEK, ITDGK 和 EYIFDR，利用 DPPH 自由
563 基清除率測定其 $IC_{50}$ 為 5.83, 3.75, 2.65, and 0.73 mM 。結果發現胜肽上具有酪
564 胺酸基者具有很好的抗自由基活性。本篇文章建議，SPLTI- a 可能有助於抗氧化
565 活性。
566

567 **關鍵詞:** 抗氧化活性; 葉子之胰蛋白腖抑制因子; 重組蛋白質; 甘藷。